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**MAR THOMA RESIDENTIAL SCHOOL ,TIRUVALLA**  
**SECOND TERMINAL EXAMINATION DEC -2019-'20**

**CHEMISTRY**

**CLASS : VIII**

**MARKS : 80**

**TIME : 2 HRS.**

**SECTION – A**

**QUESTION – 1**

**1. FILL IN THE BLANKS:**

- (a) The theory about the fundamental particles of matter was first proposed by -----, [6]
- (b) Neutron and protons present in the nucleus are called -----.
- (c) ----- is the sub atomic particle which do not have any charge.
- (d) The technique used to purify sulphide ores is called -----.
- (e) ----- are atoms to which extra electrons are added.
- (f) ----- is a mixture of metal with another metal or non-metal.

**2. DEFINE THE FOLLOWING:**

- (i) Ore (ii) Emulsion (iii) distillation (iv) valency (v) electronic configuration [5]

**3. WRITE THE FORMULAE OF THE FOLLOWING COMPOUNDS:**

- (i) Zinc sulphate (ii) Ammonium phosphate (iii) Potassium sulphide [6]
- (ii) Calcium hydroxide (v) Ferrous chloride (vi) Cupric nitrate

**4. WRITE THE ELECTRONIC CONFIGURATION:**

- (i) Potassium ( Z = 19) (ii) Lithium (Z= 3) [2]

**5. GIVE REASONS FOR THE FOLLOWING:**

- (i) An atom is always neutral. [6]
- (ii) Air is not a compound
- (iii) Ionic compounds have high melting point
- (iv) Shells of an atom are called energy levels.
- (v) Atoms combine to form compounds.
- (vi) Butter is separated from milk by centrifugation.

**6. HOW WILL YOU SEPARATE THE FOLLOWING:**

- (i) Iron filings and sulphur [5]
- (ii) Oil and water
- (iii) Alcohol and water
- (iv) Ammonium chloride and salt
- (v) Gravel from sand

**7. GIVE TWO DIFFERENCE BETWEEN THE FOLLOWING:**

- (i) Atomic number and mass number (ii) Homogeneous and heterogeneous mixture [10]

- (ii) Element and compound (iv) Electron and proton (v) alpha and beta rays

### SECTION – B

#### QUESTION – 2

(a) Mention any four points of Rutherford's model of atom.

What are the limitations of Rutherford's model of an atom.

(b) What does symbol stand for? Give the symbol and Latin name of mercury.

(c) What do you mean by atomic mass and chemical formula. How is atomic mass measured?

#### QUESTION – 3

(a) Give any four postulates of Dalton's atomic theory.

(b) What can be the valency of a metal and non-metal other than carbon?

(c) What are cathode rays? Give its charge.

(d) Draw the structure of magnesium atom. Is it a metal or non-metal. [Mg]

(e) Why are noble gases called inert gases? Name any two noble gases.

#### QUESTION – 4

(a) Give any four features of Bohr's atomic theory.

(b) Draw the geometrical structure of sodium chloride [Na = 11, Cl = 17].

State the type of bond and why is it so called?

(c) What are mixtures? Give one example each of gas in solid and solid in solid mixtures.

(d) What are suspensions? How can you separate them?

(e) What do you mean by octet rule?

#### QUESTION – 5

(a) What are the isotopes of hydrogen? How do they differ?

(b) Give any four properties of covalent compounds.

(c) Write any four characteristics of a mixture.

(d) Draw the geometrical structure of carbon dioxide [C = 6, O = 8].

State the type of bond.

(e) Classify the following as element, compound and mixture.

Blood, sugar, carbon, milk, iodine, carbon dioxide

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