

MAR THOMA RESIDENTIAL SCHOOL, TIRUVALLA
 STD VI SECOND TERMINAL EXAMINATION
 CHEMISTRY

MARK 80
 TIME 1 1/2 hrs

I. Fill in the blanks

(10)

- _____ is the method to obtain a pure solid from its solution.
- Two miscible liquids can be separated by _____.
- Changes that are useful to us are called _____.
- _____ is a liquid non metal.
- _____ increases the rate of sedimentation.
- The purity of a substance is determined by its _____ and _____.
- Molecules which contain only one atom are called _____.
- The temperature at which condensation takes place is called _____.
- The melting and freezing point of a given substance are the _____.
- A mixture of ethanol and water can be separated by _____.

II. Name the following.

- A heterogeneous solid-solid mixture.
- The process which is used to obtain common salt from sea water.
- The process of settling down of heavy particles.
- A liquid metal.
- The conditions necessary for rusting.
- A device used to separate two immiscible liquids.
- The process in which a gas turns into a liquid on cooling.
- Two sublimable substances.
- The method used to separate two immiscible liquids..
- The products of fermentation.

(10)

III. Differentiate between.

- Evaporation and distillation.
- Periodic and non-periodic changes.
- Homogeneous and heterogeneous mixtures.
- Melting and freezing.
- Exothermic and endothermic changes

(10)

IV. Define the following.

- Centrifugation
- Sieving
- Winnowing*
- Chemical formula
- Decantation.

(10)

V. Give reasons for the following.

1. Burning of magnesium is a chemical change.
2. Heating of a metal rod is a physical change.
3. Landslides and floods are undesirable changes.
4. Substances such as sugar and urea are highly pure.
5. Objects are seen more clearly after rain.
6. Burning of a candle is an example where both physical and chemical changes occur simultaneously

(12)

VI. Answer the following.

1. What is meant by rusting? Write the chemical name of rust?
2. Write the conditions which favour a chemical change?
3. Name four factors which indicate us about a chemical change.?
4. What is atomicity? Write the atomicity of fluorine.
5. How will you show that air is a mixture of gases?
6. What are the factors affecting evaporation?

(10)

VII. Classify the following in to physical and chemical changes

1. Evaporation of water
2. cutting carrots into pieces
3. Respiration
4. curdling of milk
5. Stretching of a rubber band
6. Burning of coal

(3)

VIII. Write the symbols of the following elements.

1. Mercury
2. Helium
3. Tin
4. Magnesium
5. Sodium
6. Zinc

(3)

IX. Write the chemical formula for the following

1. Hydrogen
2. Nitrogen
3. Ozone
4. Sulphur
5. Oxygen
6. Chlorine

(3)

X. Describe filtration with the help of a neat labelled diagram.

(5)

XI. Correct the following statements.

1. $2H$ represents 2 molecules of hydrogen.
2. Compounds and elements are impure substances.
3. Argon is a diatomic gas.
4. Natrium is the Latin name of potassium.

(4)